Solution Chemistry

February 5, 2007

Ionic Compound

$$\text{NaCl}_{(s)} \xrightarrow{\text{H}_2\text{O}} \text{Na}^{+}_{(aq)} + \text{Cl}^{-}_{(aq)}$$
Dissolving or Hydration

Dissolved Ions Conduct Electricity
Non-electrolyte

No Ions To Conduct Electricity

Precipitation Reaction

http://www.dlt.ncssm.edu/TIGER/Flash/moles/DoubleDisp_Reaction-Precipitation.html
### Demo

<p>| | |</p>
<table>
<thead>
<tr>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>AgNO₃</td>
</tr>
<tr>
<td>2</td>
<td>K₂Cr₂O₇</td>
</tr>
<tr>
<td>3</td>
<td>NiCl₂</td>
</tr>
<tr>
<td>4</td>
<td>Na₂CO₃</td>
</tr>
<tr>
<td>5</td>
<td>Pb(NO₃)₂</td>
</tr>
<tr>
<td>6</td>
<td>KI</td>
</tr>
<tr>
<td>7</td>
<td>NaNO₃</td>
</tr>
<tr>
<td>8</td>
<td>KCl</td>
</tr>
<tr>
<td>9</td>
<td>BaCl₂</td>
</tr>
<tr>
<td>10</td>
<td>Na₂SO₄</td>
</tr>
<tr>
<td>11</td>
<td>CuCl₂</td>
</tr>
<tr>
<td>12</td>
<td>Na₂CO₃</td>
</tr>
</tbody>
</table>

**PPT**

- Ag₂Cr₂O₇(s)
- NiCO₃(s)
- Pbl₂(s)
- No Reaction
- BaSO₄(s)
- CuCO₃(s)

### Acid-Base Reaction

[Diagram of acid-base reaction with HCl(aq) and NaOH(aq)]

http://www.dlt.ncssm.edu/TIGER/Flash/moles/DoubleDisp_Reaction-AcidToBase.html